

2,6-Dimethylpyridinium 3-carboxy-4-hydroxybenzenesulfonate monohydrate

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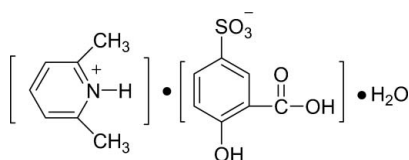
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 Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.033; wR factor = 0.110; data-to-parameter ratio = 12.5.

The asymmetric unit of the title compound, $\text{C}_7\text{H}_{10}\text{N}^+\cdot\text{C}_7\text{H}_5\text{O}_6\text{S}^-\cdot\text{H}_2\text{O}$, contains one 2,6-dimethylpyridinium cation, one 3-carboxy-4-hydroxybenzenesulfonate anion and one water molecule. Intermolecular $\text{O}-\text{H}\cdots\text{O}$, $\text{O}-\text{H}\cdots\text{S}$ and $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds, and $\pi-\pi$ interactions between the benzene and pyridine rings of neighboring ions [centroid-to-centroid distance = $3.6706(12)$ Å], generate a three-dimensional hydrogen-bonded framework structure.

Related literature

 For related literature, see: Smith *et al.* (2007).


Experimental

Crystal data

 $\text{C}_7\text{H}_{10}\text{N}^+\cdot\text{C}_7\text{H}_5\text{O}_6\text{S}^-\cdot\text{H}_2\text{O}$
 $M_r = 343.35$

 Monoclinic, $P2_1/n$
 $a = 8.8340(13)$ Å

 $b = 14.840(2)$ Å

 $c = 12.2062(17)$ Å

 $\beta = 97.057(2)^\circ$
 $V = 1588.1(4)$ Å³
 $Z = 4$

 Mo $K\alpha$ radiation

 $\mu = 0.24$ mm⁻¹
 $T = 294(2)$ K

 $0.26 \times 0.24 \times 0.20$ mm

Data collection

Bruker SMART CCD area-detector diffractometer

Absorption correction: multi-scan

 (*SADABS*; Sheldrick, 1996)

 $T_{\min} = 0.940$, $T_{\max} = 0.954$

8045 measured reflections

2810 independent reflections

 2331 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.027$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.033$
 $wR(F^2) = 0.110$
 $S = 0.99$

2810 reflections

225 parameters

4 restraints

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 0.21$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.31$ e Å⁻³
Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
O4—H4···O2 ⁱ	0.82	1.91	2.7006 (18)	160
O4—H4···S1 ⁱ	0.82	2.81	3.4464 (15)	136
N1—H1A···O7 ⁱⁱ	0.90 (1)	1.85 (1)	2.748 (2)	175 (2)
O7—H7A···O1 ⁱⁱⁱ	0.86 (1)	2.03 (1)	2.884 (2)	173 (3)
O7—H7B···O3 ^{iv}	0.86 (1)	1.93 (1)	2.787 (2)	173 (3)
O6—H6···O5	0.82	1.93	2.647 (2)	145

 Symmetry codes: (i) $-x, -y + 2, -z + 2$; (ii) $-x + 2, -y + 1, -z + 1$; (iii) $-x + 1, -y + 1, -z + 1$; (iv) $x + 1, y - 1, z$.

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINTE* (Bruker, 2000); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *SHELXTL* (Bruker, 2001); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2511).

References

- Bruker (2000). *SMART* (Version 5.051) and *SAINTE* (Version 5. A06). Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2001). *SHELXTL*. Version 6.12. Bruker AXS Inc., Madison, Wisconsin, USA.
- Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Smith, G., Wermuth, U. D., Young, D. J. & White, J. M. (2007). *Polyhedron*, **26**, 3645–3652.

supplementary materials

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2,6-Dimethylpyridinium 3-carboxy-4-hydroxybenzenesulfonate monohydrate

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Comment

5-Sulfosalicylic acid (SSA) has six potential donor sites in the three substituent groups (the sulfonic acid, the carboxylic acid and the phenolic groups), and it gives mono-, di- and trianionic ligand species through deprotonation. The presence of numerous oxygen atoms in the substituent groups usually results in hydrogen-bonding associations, and the self-assembly process of crystallization often requires the incorporation of water molecules in the structures (Smith *et al.* 2007). We report here the crystal structure of the title compound.

The asymmetric unit of the title compound contains one 2,6-dimethyl pyridinium cation, one 3-carboxyl-4-hydroxy-benzenesulfonate anion and one water molecule (Fig. 1). The bond distances and angles in the cationic and anionic species are normal. An intramolecular O6—H6···O5 hydrogen bond is observed. The molecular packing (Fig. 2) is stabilized by intermolecular O—H···O, O—H···S and N—H···O hydrogen bonds (Table 1), and π - π interactions between the benzene and pyridine rings of the neighboring ions [centroid-to-centroid distance is 3.6706 (12) Å]. These interactions generate a three-dimensional hydrogen-bonded framework structure.

Experimental

2-Hydroxy-5-sulfobenzoic acid (2.18 g, 10 mmol), 2,6-dimethylpyridine (1.07 g, 10 mmol) and H₂O (20 ml) were loaded into a 50 ml roundbottom flask, and heated to dissolve the solid. Crystals of the title compound were obtained by slow evaporation of the deionic H₂O solution.

Refinement

The H atoms of the water molecule, and the N-bound H atom were located in a difference Fourier map, and refined with the O—H and N—H distance restraints of 0.86 (1) and 0.90 (1) Å, respectively. All other H atoms were positioned geometrically [O—H = 0.82 Å (hydroxyl), C—H = 0.93 Å (aromatic) and 0.96 Å (methyl)] and refined using a riding model, with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{carrier})$ for hydroxyl and methyl H atoms and $1.2U_{\text{eq}}(\text{C})$ for other H atoms.

Figures

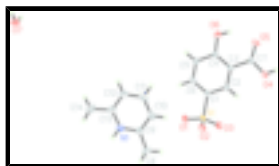


Fig. 1. The asymmetric unit of title compound. Displacement ellipsoids are drawn at the 30% probability level.

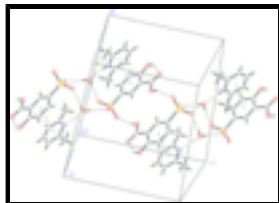


Fig. 2. Part of the crystal packing of the title compound. π - π interactions, O—H...O and N—H...O hydrogen bonds are shown as dashed lines.

2,6-Dimethylpyridinium 3-carboxy-4-hydroxybenzenesulfonate monohydrate

Crystal data

$C_7H_{10}N^+ \cdot C_7H_5O_6S^- \cdot H_2O$

$M_r = 343.35$

Monoclinic, $P2_1/n$

Hall symbol: -P 2yn

$a = 8.8340$ (13) Å

$b = 14.840$ (2) Å

$c = 12.2062$ (17) Å

$\beta = 97.057$ (2)°

$V = 1588.1$ (4) Å³

$Z = 4$

$F_{000} = 720$

$D_x = 1.436$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 4380 reflections

$\theta = 2.7$ – 26.4 °

$\mu = 0.24$ mm⁻¹

$T = 294$ (2) K

Block, colourless

$0.26 \times 0.24 \times 0.20$ mm

Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 294$ (2) K

φ and ω scans

Absorption correction: multi-scan (SADABS; Sheldrick, 1996)

$T_{\min} = 0.940$, $T_{\max} = 0.954$

8045 measured reflections

2810 independent reflections

2331 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.027$

$\theta_{\text{max}} = 25.0$ °

$\theta_{\text{min}} = 2.2$ °

$h = -9 \rightarrow 10$

$k = -17 \rightarrow 11$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.033$

$wR(F^2) = 0.110$

$S = 0.99$

2810 reflections

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0735P)^2 + 0.3819P]$$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\text{max}} = 0.003$

$\Delta\rho_{\text{max}} = 0.21$ e Å⁻³

225 parameters

$$\Delta\rho_{\min} = -0.31 \text{ e } \text{\AA}^{-3}$$

4 restraints

Extinction correction: SHELXL97,
 $F_c^* = kF_c [1 + 0.001 \times F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$

Primary atom site location: structure-invariant direct methods

Extinction coefficient: 0.088 (5)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.19980 (5)	0.97070 (3)	0.73325 (3)	0.03601 (19)
O1	0.24931 (16)	0.93374 (10)	0.63251 (11)	0.0483 (4)
O2	0.32890 (14)	0.99904 (9)	0.81389 (11)	0.0456 (4)
O3	0.08412 (16)	1.04152 (9)	0.71330 (12)	0.0515 (4)
O4	-0.21320 (16)	0.93634 (9)	1.00700 (12)	0.0481 (4)
H4	-0.2600	0.9446	1.0601	0.072*
O5	-0.23614 (17)	0.79077 (9)	1.05258 (12)	0.0513 (4)
O6	-0.08353 (17)	0.66773 (9)	0.95369 (12)	0.0507 (4)
H6	-0.1434	0.6857	0.9951	0.076*
N1	0.76930 (17)	0.80636 (10)	0.61307 (13)	0.0399 (4)
C1	0.10902 (18)	0.88043 (11)	0.79670 (13)	0.0309 (4)
C2	0.00140 (18)	0.89931 (11)	0.86792 (13)	0.0308 (4)
H2	-0.0255	0.9588	0.8797	0.037*
C3	-0.06686 (18)	0.82946 (11)	0.92206 (13)	0.0310 (4)
C4	-0.0247 (2)	0.73902 (11)	0.90409 (14)	0.0351 (4)
C5	0.0842 (2)	0.72112 (12)	0.83225 (15)	0.0407 (4)
H5	0.1127	0.6619	0.8205	0.049*
C6	0.1497 (2)	0.79065 (12)	0.77869 (14)	0.0372 (4)
H6A	0.2208	0.7778	0.7307	0.045*
C7	-0.18011 (19)	0.84923 (12)	0.99981 (14)	0.0347 (4)
C8	0.6895 (3)	0.96258 (14)	0.6310 (2)	0.0643 (6)
H8A	0.7911	0.9842	0.6532	0.096*
H8B	0.6186	0.9965	0.6683	0.096*
H8C	0.6650	0.9697	0.5527	0.096*
C9	0.6800 (2)	0.86485 (13)	0.66063 (16)	0.0434 (5)
C10	0.5833 (2)	0.83132 (16)	0.73238 (19)	0.0551 (6)
H10	0.5226	0.8701	0.7680	0.066*
C11	0.5786 (2)	0.73910 (18)	0.7501 (2)	0.0645 (7)

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H11	0.5138	0.7158	0.7976	0.077*
C12	0.6690 (3)	0.68186 (16)	0.6981 (2)	0.0599 (6)
H12	0.6635	0.6201	0.7096	0.072*
C13	0.7683 (2)	0.71537 (13)	0.62877 (16)	0.0449 (5)
C14	0.8735 (3)	0.65944 (15)	0.5681 (2)	0.0632 (6)
H14A	0.8272	0.6493	0.4939	0.095*
H14B	0.8921	0.6027	0.6050	0.095*
H14C	0.9684	0.6908	0.5669	0.095*
O7	1.01536 (19)	0.12730 (11)	0.51013 (13)	0.0578 (4)
H1A	0.836 (2)	0.8274 (15)	0.5687 (15)	0.059 (7)*
H7A	0.9323 (19)	0.1105 (17)	0.4717 (16)	0.079 (9)*
H7B	1.038 (3)	0.0964 (16)	0.5696 (14)	0.079 (9)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S1	0.0364 (3)	0.0391 (3)	0.0354 (3)	0.00222 (17)	0.01625 (19)	0.00305 (18)
O1	0.0522 (8)	0.0591 (9)	0.0381 (7)	0.0007 (7)	0.0232 (6)	-0.0004 (6)
O2	0.0414 (7)	0.0505 (8)	0.0474 (8)	-0.0074 (6)	0.0152 (6)	-0.0069 (6)
O3	0.0534 (8)	0.0476 (8)	0.0579 (9)	0.0147 (6)	0.0241 (7)	0.0182 (7)
O4	0.0560 (9)	0.0419 (8)	0.0524 (8)	-0.0009 (6)	0.0314 (7)	-0.0035 (6)
O5	0.0579 (9)	0.0465 (8)	0.0549 (8)	-0.0069 (7)	0.0294 (7)	0.0069 (7)
O6	0.0649 (9)	0.0343 (7)	0.0563 (9)	-0.0062 (6)	0.0211 (7)	0.0038 (6)
N1	0.0381 (8)	0.0388 (9)	0.0415 (9)	-0.0063 (7)	0.0002 (7)	0.0063 (7)
C1	0.0293 (8)	0.0371 (9)	0.0269 (8)	0.0008 (7)	0.0061 (7)	-0.0006 (7)
C2	0.0314 (8)	0.0315 (9)	0.0303 (9)	0.0011 (7)	0.0063 (7)	-0.0008 (7)
C3	0.0299 (8)	0.0346 (9)	0.0284 (8)	-0.0019 (7)	0.0031 (7)	-0.0013 (7)
C4	0.0378 (9)	0.0330 (9)	0.0340 (9)	-0.0043 (7)	0.0025 (7)	0.0002 (7)
C5	0.0471 (10)	0.0314 (9)	0.0440 (10)	0.0050 (8)	0.0078 (8)	-0.0056 (8)
C6	0.0359 (9)	0.0428 (10)	0.0340 (9)	0.0048 (8)	0.0093 (7)	-0.0063 (8)
C7	0.0326 (9)	0.0386 (9)	0.0338 (9)	-0.0041 (7)	0.0076 (7)	-0.0009 (7)
C8	0.0694 (15)	0.0409 (12)	0.0827 (17)	0.0032 (10)	0.0094 (13)	0.0018 (11)
C9	0.0366 (10)	0.0448 (11)	0.0473 (11)	-0.0013 (8)	-0.0017 (8)	0.0038 (8)
C10	0.0365 (10)	0.0680 (15)	0.0606 (13)	0.0019 (10)	0.0057 (9)	0.0082 (11)
C11	0.0412 (11)	0.0811 (17)	0.0718 (15)	-0.0072 (11)	0.0091 (11)	0.0309 (13)
C12	0.0489 (12)	0.0477 (12)	0.0794 (16)	-0.0120 (10)	-0.0066 (11)	0.0247 (11)
C13	0.0410 (10)	0.0384 (10)	0.0511 (11)	-0.0048 (8)	-0.0116 (9)	0.0044 (9)
C14	0.0650 (14)	0.0471 (12)	0.0733 (15)	0.0041 (11)	-0.0087 (12)	-0.0101 (11)
O7	0.0619 (10)	0.0654 (10)	0.0469 (9)	-0.0062 (8)	0.0093 (8)	0.0132 (8)

Geometric parameters (\AA , $^\circ$)

S1—O1	1.4616 (13)	C5—H5	0.93
S1—O3	1.4652 (14)	C6—H6A	0.93
S1—O2	1.4733 (14)	C8—C9	1.500 (3)
S1—C1	1.7861 (17)	C8—H8A	0.96
O4—C7	1.331 (2)	C8—H8B	0.96
O4—H4	0.82	C8—H8C	0.96
O5—C7	1.221 (2)	C9—C10	1.388 (3)

O6—C4	1.354 (2)	C10—C11	1.387 (3)
O6—H6	0.82	C10—H10	0.93
N1—C9	1.351 (3)	C11—C12	1.374 (4)
N1—C13	1.364 (2)	C11—H11	0.93
N1—H1A	0.904 (10)	C12—C13	1.384 (3)
C1—C2	1.393 (2)	C12—H12	0.93
C1—C6	1.404 (2)	C13—C14	1.506 (3)
C2—C3	1.405 (2)	C14—H14A	0.96
C2—H2	0.93	C14—H14B	0.96
C3—C4	1.417 (2)	C14—H14C	0.96
C3—C7	1.490 (2)	O7—H7A	0.858 (9)
C4—C5	1.405 (3)	O7—H7B	0.861 (9)
C5—C6	1.385 (3)		
O1—S1—O3	113.83 (8)	O5—C7—C3	122.91 (16)
O1—S1—O2	112.46 (8)	O4—C7—C3	113.79 (15)
O3—S1—O2	111.91 (9)	C9—C8—H8A	109.5
O1—S1—C1	106.26 (8)	C9—C8—H8B	109.5
O3—S1—C1	105.60 (8)	H8A—C8—H8B	109.5
O2—S1—C1	106.04 (8)	C9—C8—H8C	109.5
C7—O4—H4	109.5	H8A—C8—H8C	109.5
C4—O6—H6	109.5	H8B—C8—H8C	109.5
C9—N1—C13	124.16 (17)	N1—C9—C10	118.54 (19)
C9—N1—H1A	119.6 (15)	N1—C9—C8	117.70 (19)
C13—N1—H1A	116.2 (15)	C10—C9—C8	123.7 (2)
C2—C1—C6	119.71 (16)	C11—C10—C9	119.0 (2)
C2—C1—S1	119.80 (13)	C11—C10—H10	120.5
C6—C1—S1	120.44 (13)	C9—C10—H10	120.5
C1—C2—C3	120.70 (16)	C12—C11—C10	120.5 (2)
C1—C2—H2	119.6	C12—C11—H11	119.7
C3—C2—H2	119.6	C10—C11—H11	119.7
C2—C3—C4	119.34 (15)	C11—C12—C13	120.6 (2)
C2—C3—C7	121.00 (15)	C11—C12—H12	119.7
C4—C3—C7	119.65 (15)	C13—C12—H12	119.7
O6—C4—C5	117.47 (16)	N1—C13—C12	117.2 (2)
O6—C4—C3	123.25 (16)	N1—C13—C14	117.49 (19)
C5—C4—C3	119.27 (16)	C12—C13—C14	125.3 (2)
C6—C5—C4	120.76 (16)	C13—C14—H14A	109.5
C6—C5—H5	119.6	C13—C14—H14B	109.5
C4—C5—H5	119.6	H14A—C14—H14B	109.5
C5—C6—C1	120.21 (16)	C13—C14—H14C	109.5
C5—C6—H6A	119.9	H14A—C14—H14C	109.5
C1—C6—H6A	119.9	H14B—C14—H14C	109.5
O5—C7—O4	123.29 (16)	H7A—O7—H7B	113.5 (15)
O1—S1—C1—C2	-155.64 (13)	C2—C1—C6—C5	-0.6 (3)
O3—S1—C1—C2	-34.41 (16)	S1—C1—C6—C5	176.90 (13)
O2—S1—C1—C2	84.50 (15)	C2—C3—C7—O5	-176.23 (16)
O1—S1—C1—C6	26.86 (16)	C4—C3—C7—O5	2.4 (3)
O3—S1—C1—C6	148.08 (15)	C2—C3—C7—O4	3.2 (2)

supplementary materials

O2—S1—C1—C6	-93.01 (15)	C4—C3—C7—O4	-178.16 (15)
C6—C1—C2—C3	0.1 (2)	C13—N1—C9—C10	1.7 (3)
S1—C1—C2—C3	-177.46 (12)	C13—N1—C9—C8	-177.32 (18)
C1—C2—C3—C4	0.3 (2)	N1—C9—C10—C11	-1.8 (3)
C1—C2—C3—C7	178.96 (15)	C8—C9—C10—C11	177.1 (2)
C2—C3—C4—O6	179.29 (16)	C9—C10—C11—C12	0.4 (3)
C7—C3—C4—O6	0.6 (2)	C10—C11—C12—C13	1.2 (4)
C2—C3—C4—C5	-0.1 (2)	C9—N1—C13—C12	-0.1 (3)
C7—C3—C4—C5	-178.82 (16)	C9—N1—C13—C14	179.10 (17)
O6—C4—C5—C6	-179.86 (16)	C11—C12—C13—N1	-1.4 (3)
C3—C4—C5—C6	-0.4 (3)	C11—C12—C13—C14	179.5 (2)
C4—C5—C6—C1	0.8 (3)		

Hydrogen-bond geometry (\AA , $^\circ$)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
O4—H4 \cdots O2 ⁱ	0.82	1.91	2.7006 (18)	160
O4—H4 \cdots S1 ⁱ	0.82	2.81	3.4464 (15)	136
N1—H1A \cdots O7 ⁱⁱ	0.90 (1)	1.85 (1)	2.748 (2)	175 (2)
O7—H7A \cdots O1 ⁱⁱⁱ	0.86 (1)	2.03 (1)	2.884 (2)	173 (3)
O7—H7B \cdots O3 ^{iv}	0.86 (1)	1.93 (1)	2.787 (2)	173 (3)
O6—H6 \cdots O5	0.82	1.93	2.647 (2)	145

Symmetry codes: (i) $-x, -y+2, -z+2$; (ii) $-x+2, -y+1, -z+1$; (iii) $-x+1, -y+1, -z+1$; (iv) $x+1, y-1, z$.

Fig. 1

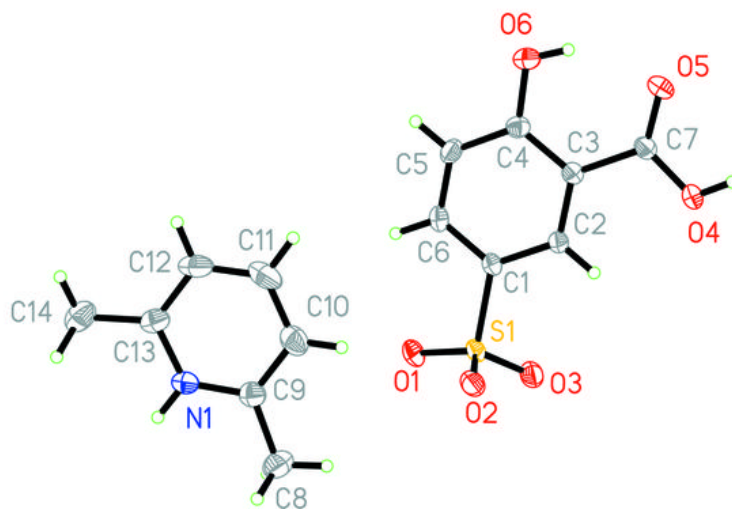


Fig. 2

